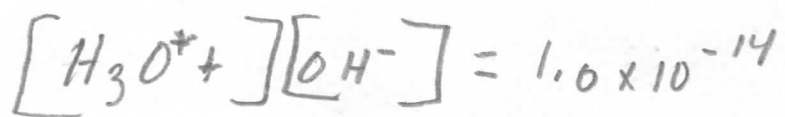


$$K_w = 1.0 \times 10^{-14}$$



$$[H_3O^+] = [OH^-] = x$$

$$x^2 = 1 \times 10^{-14}$$

$$x = 1.0 \times 10^{-7} \frac{\text{mole}}{\text{litre}} = [H_3O^+] = [OH^-]$$